

**The Crystal Structure of Two Modifications of
Chromium(III) Tris(diethyldithiophosphate),
 $\text{Cr}[\text{S}_2\text{P}(\text{OC}_2\text{H}_5)_2]_3$**

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Chromium(III) tris(diethyldithiophosphate), $\text{Cr}[\text{S}_2\text{P}(\text{OC}_2\text{H}_5)_2]_3$ crystallises in two different forms, one of which is isomorphous to the corresponding vanadium complex. The structure of the other crystal modification has been determined from three-dimensional X-ray data collected on an automatic diffractometer and refined by the method of least squares to give an *R*-value of 8.1 % for 1368 reflections. The crystal is monoclinic with $a = 14.28 \text{ \AA}$, $b = 13.60 \text{ \AA}$, $c = 14.35 \text{ \AA}$, $\beta = 89.5^\circ$; space group *C2/c*. The relatively small number of observed reflections is due to large thermal vibration and limits the accuracy of the structure determination. The molecule nearly has the symmetry 32 although only a twofold axis is required by the space group. The only difference between the two modifications is in the packing.

Jørgensen has prepared several diethyldithiophosphate (dtp) complexes¹ and kindly sent us samples of some. Those of chromium, cobalt, ruthenium, rhodium, and iridium were found to be isomorphous as indicated by their powder diagrams and by the fact that all of these complexes can co-crystallise. Lacroix has measured the paramagnetic resonance of Crdtp_3 as a 1 % impurity in Codtp_3 .² Crdtp_3 was chosen for the structure determination for this reason and because plenty of good crystals were available.

After this structure had been solved the structures of Vdtp_3 ³ and Indtp_3 ⁴ have been published. They are mutually isomorphous but different from the structure found in the present investigation. Careful investigation of our sample of Crdtp_3 under the polarising microscope showed that about 10 % of the crystals extinguished parallel to their prominent direction. These crystals were found to have cell dimensions and space group similar to those reported for Vdtp_3 . The quality of the photographs was low, but we consider it proved that this modification of Crdtp_3 is isomorphous to Vdtp_3 , and we intend to do no further work on these crystals.

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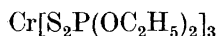
EXPERIMENTAL

The unit cell and space group were determined from precession photographs taken with $\text{MoK}\alpha$ radiation ($\lambda = 0.7107 \text{ \AA}$). Preliminary intensities for two projections were obtained from integrated precession films. The three-dimensional intensities were measured on a linear diffractometer of the Arndt-Phillips type⁶ with Mo radiation. The crystal was approximately $0.3 \times 0.3 \times 0.2 \text{ mm}^3$. Balanced filters and pulse height discrimination were applied. The absorption coefficient is $\mu = 10.7 \text{ cm}^{-1}$ for $\text{MoK}\alpha$ radiation, so it was considered justified not to correct for absorption.

STRUCTURE DETERMINATION

The positions of the heavier atoms were found from the bc and ac projections by aid of the characteristic intensity distributions. For some of the light atoms only approximate coordinates could be obtained because of overlap and lack of data. When three-dimensional data became available the trial structure gave $R = 31 \%$ and a Fourier map showed that one carbon atom had to be moved 1 \AA . Least squares refinement lowered R to 14% and further refinement including anisotropic temperature factor parameters lowered R to 9% . At this point a difference Fourier map was calculated which showed no extra atoms heavier than hydrogen. The hydrogen atoms of the CH_2 groups were found in approximately the positions expected. Those of the CH_3 groups could not be found with certainty; one would expect a staggered configuration around the $\text{C}-\text{C}$ bond but this does not correspond very well to the peaks in the difference map although there are peaks at approximately the right distances from the carbon atoms. One of the three independent ethyl groups shows large temperature factors and an abnormally short $\text{C}-\text{C}$ bond (1.34 \AA). The difference Fourier map gave no indication of the reason for this, and an attempt to move the atoms to the commonly accepted distances failed because least squares refinement immediately brought them back. A riding motion correction to the bond lengths of this OCH_2CH_3 group indicated that the $\text{O}-\text{C}$ bond appears shortened by 0.04 \AA but since both carbon atoms move nearly equally much, little correction was indicated for the $\text{C}-\text{C}$ bond. The assumption that the carbon atoms moved independently led to a correction of 0.33 \AA , so it seems likely that a model with partly correlated motion could lead to a corrected distance of the length of a normal single bond. The reason for the unusually large thermal motion of this ethyl group seems to be the lack of normal van der Waals contacts.

CRYSTAL DATA



Crystal system: monoclinic

$a = 14.28 \pm 0.03 \text{ \AA}$, $b = 13.60 \pm 0.03 \text{ \AA}$, $c = 14.35 \pm 0.03 \text{ \AA}$, $\beta = 89.5 \pm 0.02^\circ$.

Systematic absences: hkl : $h + k \neq 2n$

$h0l$: $l \neq 2n$

$d_{\text{meas}} = 1.44 \text{ g/cm}^3$, $d_{\text{calc}} = 1.45 \text{ g/cm}^3$ for four molecules per cell.

Possible space groups: *C2/c* (No. 15) and *Cc* (No. 9). No piezoelectric effect could be detected so the centrosymmetric space group *C2/c* was preferred, and the structure analysis confirmed this.

The final atomic coordinates are given in Table 1, temperature factor parameters in Table 2, interatomic distances and angles in Table 3, and Table 4 is a list of observed and calculated structure factors. The scattering factors used were those of International Tables for X-Ray Crystallography, Vol. 3, pp. 202–205.

Table 1. Atomic coordinates in fractions of the unit cell. Standard deviations $\times 10^{-4}$ in parentheses.

	<i>x</i>	$\sigma(x)$	<i>y</i>	$\sigma(y)$	<i>z</i>	$\sigma(z)$
Cr	0.0000	(0)	0.3185	(2)	0.2500	(0)
S ₁	0.0820	(2)	0.3348	(2)	0.1023	(2)
S ₂	0.1162	(2)	0.4376	(2)	0.2996	(2)
S ₃	0.0907	(2)	0.1845	(2)	0.3139	(2)
P ₁	0.0000	(0)	0.0963	(3)	0.2500	(0)
P ₂	0.1657	(2)	0.4264	(2)	0.1692	(2)
O ₁	0.0477	(6)	0.0184	(5)	0.1839	(5)
O ₂	0.1789	(5)	0.5277	(6)	0.1185	(6)
O ₃	0.2281	(5)	0.8962	(7)	0.3386	(7)
C ₁	0.1152	(12)	0.0482	(9)	0.1155	(11)
C ₂	0.1384	(11)	0.0307	(12)	0.0550	(11)
C ₃	0.1035	(9)	0.5900	(9)	0.0958	(10)
C ₄	0.1411	(10)	0.6807	(9)	0.0593	(9)
C ₅	0.1961	(12)	0.8046	(14)	0.3105	(15)
C ₆	0.1128	(15)	0.7783	(13)	0.3450	(13)

Table 2. Mean square vibration amplitudes, u_{ij} , in $\text{Å}^2 \times 10^{-4}$ with standard deviations in parentheses.

	u_{11}	σu_{11}	u_{22}	σu_{22}	u_{33}	σu_{33}	u_{12}	σu_{12}	u_{13}	σu_{13}	u_{23}	σu_{23}
Cr	603	(13)	551	(13)	769	(12)	0	(0)	19	(11)	0	(0)
S ₁	866	(20)	816	(20)	802	(14)	9	(16)	138	(16)	29	(16)
S ₂	789	(19)	796	(20)	986	(18)	-188	(16)	-56	(17)	38	(17)
S ₃	840	(20)	677	(18)	971	(17)	83	(15)	-31	(16)	84	(16)
P ₁	878	(30)	601	(25)	989	(26)	0	(0)	280	(25)	0	(0)
P ₂	678	(18)	769	(20)	1193	(22)	21	(16)	102	(18)	159	(19)
O ₁	1103	(55)	700	(47)	1210	(49)	-33	(43)	426	(48)	-130	(45)
O ₂	858	(51)	835	(53)	1520	(59)	-78	(43)	108	(49)	334	(50)
O ₃	671	(49)	1200	(69)	2066	(77)	-89	(49)	-91	(56)	-358	(64)
C ₁	1664	(123)	847	(87)	1679	(113)	-5	(87)	1038	(105)	-186	(86)
C ₂	1531	(126)	1309	(121)	1649	(106)	307	(105)	712	(104)	11	(102)
C ₃	942	(86)	768	(78)	1727	(105)	109	(68)	-5	(85)	378	(80)
C ₄	1138	(99)	784	(84)	1483	(99)	-38	(74)	-160	(87)	360	(81)
C ₅	1011	(103)	1939	(158)	2184	(146)	-638	(107)	178	(111)	-671	(131)
C ₆	2007	(165)	1615	(142)	2114	(152)	-660	(134)	964	(139)	-88	(129)

Table 3. Interatomic distances and angles with standard deviations in parentheses.

Distance	Å	10 ⁻³ Å	Distance	Å	10 ⁻³ Å
Cr-S ₁	2.423	(3)	P ₂ -O ₃	1.574	(9)
Cr-S ₂	2.430	(3)	O ₁ -C ₁	1.429	(17)
Cr-S ₃	2.421	(3)	O ₂ -C ₃	1.412	(15)
S ₁ -P ₂	1.980	(4)	O ₃ -C ₅	1.388	(22)
S ₂ -P ₂	2.000	(5)	C ₁ -C ₂	1.418	(21)
S ₃ -P ₁	1.995	(4)	C ₃ -C ₄	1.441	(18)
P ₁ -O ₁	1.572	(8)	C ₅ -C ₆	1.334	(27)
P ₂ -O ₂	1.569	(9)			

Angle	Degrees		Angle	Degrees	
S ₁ -Cr-S ₂	82.5	(0.1)	S ₁ -P ₂ -O ₂	113.5	(0.4)
S ₁ -Cr-S ₃	98.3	(0.1)	S ₁ -P ₂ -O ₃	112.8	(0.4)
S ₂ -Cr-S ₃	91.3	(0.1)	S ₂ -P ₁ -O ₂	114.1	(0.4)
S ₃ -Cr-S ₃	82.4	(0.1)	S ₂ -P ₂ -O ₃	114.7	(0.4)
Cr-S ₁ -P ₂	85.5	(0.1)	O ₂ -P ₂ -O ₃	94.8	(0.5)
Cr-S ₂ -P ₂	84.9	(0.1)	P ₁ -O ₁ -C ₁	120.6	(0.7)
Cr-S ₃ -P ₁	85.9	(0.1)	P ₂ -O ₂ -C ₃	123.1	(0.8)
S ₂ -P ₁ -S ₃	106.2	(0.2)	P ₂ -O ₃ -C ₅	122.2	(1.0)
S ₃ -P ₁ -O ₁	113.8	(0.3)	O ₁ -C ₁ -C ₂	111.0	(1.1)
O ₁ -P ₁ -O ₁	95.3	(0.4)	O ₂ -C ₃ -C ₄	108.4	(1.1)
S ₁ -P ₂ -S ₂	107.0	(0.2)	O ₃ -C ₅ -C ₆	115.4	(1.6)

DESCRIPTION ON THE STRUCTURE

The molecule could be expected to have 32 symmetry and indeed this is nearly the case although only one twofold axis is imposed by the space group.

The chromium atoms are surrounded by six sulphur atoms in a slightly flattened octahedron, the angles at chromium in the Cr-S-P-S ring being only 82.3°. The Cr-S distances are in good agreement with those found in other compounds, *e.g.* 2.42-2.46 in the chromium sulphides.⁶ The average S-P distance is 1.99 Å which is in between the S-P distance in SP(C₂H₅)₃ (1.86 Å) (double bond),⁷ and the calculated length of a single bond (2.10 Å).⁸ It is close to the value (2.01 Å) found in bis(dimethyl dithiophosphinato)-nickel⁹ and to the average of those in Vdtp₃³ (1.98 Å), in Indtp₃⁴ (1.99 Å), and in Bidtp₃¹⁰ (1.99 Å). No indication is found for two different S-P bond lengths as reported for Indtp₃.⁴ The P-O distances are close to the length of the P-O single bond as found in the ferro-electric form of KH₂PO₄¹¹ so the double bond character of the thiophosphate group seems distributed between the two P-S bonds.

The S-P-S angles are fairly close to the tetrahedral value; they have to be large so that the ligand can span the Cr-ion. The Cr-S-P angles are slightly smaller than the value of 95-100°, usually found for angles of similar sulphur atoms. The positions of oxygen and carbon atoms are of low accuracy, but the distances and angles between them do not differ significantly from the

Table 4. Observed and calculated structure factors, 10 × absolute values.

h	k	l	Prob	Calc	h	k	l	Prob	Calc	h	k	l	Prob	Calc	h	k	l	Prob	Calc
-1	0	0	250	258	6	10	1	571	-596	-12	2	1	136	111	-6	6	2	283	287
-1	0	0	201	-236	5	11	1	182	-159	-14	1	1	226	-235	-7	7	2	190	210
0	12	0	422	466	4	12	1	507	525	-13	1	1	263	-272	-9	9	2	425	-466
-1	0	0	204	-301	3	13	1	395	-410	13	1	2	169	115	-2	0	2	445	-409
-2	14	0	237	-256	15	2	1	973	289	15	5	2	162	237	1	1	1	601	-620
-4	16	0	119	-67	12	2	1	143	85	10	10	2	150	-182	-5	3	2	486	-468
-1	1	0	123	-47	11	3	1	141	171	15	3	2	236	-232	-6	4	2	765	826
-2	12	0	189	212	10	4	1	283	-214	14	4	2	134	100	-1	5	2	907	932
-3	12	0	292	246	8	6	1	335	-232	13	5	2	162	101	-8	6	2	167	-158
-4	14	0	170	-163	7	7	1	282	286	10	8	2	214	205	-9	7	2	143	150
-7	17	0	163	-79	4	10	1	122	-149	9	9	2	146	-46	-10	8	2	398	820
0	8	0	1861	-1360	3	11	1	207	-307	8	10	2	170	-175	-11	9	2	290	-279
-1	9	0	830	870	1	13	1	244	-231	6	12	2	178	-172	-4	0	2	950	-915
-2	10	0	202	199	-1	15	1	199	192	4	14	2	300	222	-5	1	2	107	-111
-5	13	0	250	246	12	1	1	457	-479	3	15	2	156	-112	-6	2	2	312	275
-1	0	0	901	865	10	2	1	441	-454	13	5	2	143	-152	-7	2	1	1075	-1266
-1	7	0	296	286	8	4	1	545	-573	12	4	2	146	161	-8	4	2	437	456
-2	8	0	445	-446	7	5	1	566	-546	12	5	2	108	221	-9	5	2	794	795
-4	10	0	827	430	6	6	1	487	540	10	6	2	399	-205	-10	6	2	146	-166
-5	11	0	178	-206	5	7	1	276	259	8	8	2	512	518	-14	10	2	154	-159
0	4	0	429	-344	4	8	1	542	569	7	9	2	504	-524	-6	0	2	413	-367
-1	5	0	275	256	2	9	1	363	367	5	10	2	307	326	-7	1	2	322	281
-2	6	0	1079	1056	10	2	1	342	-404	6	12	2	182	-194	-8	2	1	395	407
-3	7	0	503	-504	11	3	1	370	-304	3	13	2	153	-200	-9	3	2	304	-370
-4	8	0	731	724	8	4	1	403	-461	2	14	2	177	193	-10	4	2	224	245
-5	9	0	156	-162	-2	13	1	287	-272	6	15	2	145	-159	-11	5	2	146	-166
-6	10	0	194	184	9	1	1	161	-191	5	16	2	198	-156	-8	0	2	148	115
-7	11	0	267	-261	8	2	1	201	-203	13	1	2	143	141	-9	1	2	552	547
0	2	0	758	767	5	3	1	387	-391	8	6	2	267	272	-11	3	2	126	8
-1	3	0	1035	959	4	4	1	614	-636	6	7	2	405	-419	-14	4	2	166	164
-2	4	0	270	-263	3	5	1	295	-309	6	8	2	295	303	-10	0	2	1237	-1266
-3	5	0	1187	-1166	3	7	1	506	-535	5	9	2	363	-426	-11	2	2	355	379
-4	6	0	225	262	-1	8	1	353	367	1	10	2	220	-247	-12	2	2	182	178
-5	7	0	948	-955	5	9	1	716	-738	3	11	2	510	519	-12	0	2	301	-303
-7	9	0	213	223	0	10	1	654	-640	2	12	2	261	-304	-13	1	2	351	-337
-8	10	0	121	111	-5	13	1	330	-322	-1	15	2	175	152	-14	2	2	137	186
-9	11	0	738	-733	-4	14	1	475	-475	9	3	1	342	-347	-15	3	2	184	-184
-10	12	0	820	-858	-5	15	1	143	-144	11	1	2	562	-340	-13	1	3	151	-27
-11	13	0	1403	1365	1	16	1	529	-528	10	2	2	521	579	-12	2	3	142	-104
-12	14	0	271	-257	6	17	1	353	357	9	3	2	306	273	-11	3	3	175	175
-13	15	0	1468	-1455	4	18	1	272	-273	8	4	2	405	-416	-15	4	3	233	224
-6	6	0	524	-549	3	19	1	1123	1127	5	7	2	290	318	-14	6	3	145	-165
-7	7	0	410	443	-1	20	1	831	-831	4	8	2	516	-509	-10	10	3	191	-175
-8	8	0	250	-246	-4	21	1	608	-628	2	10	2	405	-416	-14	8	3	150	-155
-9	9	0	183	210	-2	22	1	135	130	2	10	2	139	-155	9	3	3	307	-311
-10	10	0	17	93	-3	23	1	104	-104	1	11	2	164	146	-10	10	3	271	329
-11	12	0	170	131	-5	24	1	286	-289	0	12	2	154	-165	-7	1	3	215	-237
-12	13	0	1478	-1468	1	25	1	347	-347	1	13	2	205	187	-11	2	3	142	-142
-13	14	0	1097	1024	3	3	1	104	1157	10	0	2	151	49	-14	2	3	169	-114
-4	2	0	567	-618	2	4	1	1959	-1834	9	1	2	309	-332	-12	4	3	202	208
-5	3	0	208	-214	1	5	1	608	-628	11	8	2	205	-232	-5	5	3	192	-198
-6	4	0	189	-187	0	6	1	232	262	7	3	2	265	239	9	3	3	304	-306
-7	5	0	316	-365	-1	7	1	1405	-1472	5	5	2	995	955	-7	3	3	539	-561
-8	6	0	352	-344	-2	8	1	395	-435	4	6	2	816	-826	-6	10	3	405	414
-9	7	0	205	-199	-1	9	1	280	-280	3	7	2	648	-648	-11	13	3	242	-249
-10	8	0	391	-303	-6	12	1	244	238	2	8	2	245	262	-6	13	3	253	215
-12	10	0	138	77	3	1	1	975	-997	1	9	2	596	-514	-14	15	3	158	15
-13	11	0	2279	-2298	1	2	1	104	-109	0	10	2	503	-302	-10	16	3	144	-9
-1	0	0	344	-375	1	3	1	1250	-1318	1	1	2	938	987	-11	17	3	203	224
-2	1	0	748	752	0	4	1	150	-157	-1	3	2	201	-215	-12	2	3	145	118
-3	2	0	424	446	-1	5	1	30	-107	-4	4	2	175	195	-11	3	3	333	-360
-4	3	0	263	-268	-1	6	1	608	-628	1	5	2	151	-161	-10	4	3	182	-182
-5	4	0	405	-416	-4	7	1	153	-133	7	1	2	978	972	9	5	3	134	-135
-6	5	0	270	287	-5	8	1	346	373	6	2	2	187	-157	8	6	3	155	-179
-7	6	0	102	-109	-6	10	1	309	-319	5	3	2	1952	-1789	7	7	3	627	638
-8	7	0	240	-238	-7	11	1	650	-650	4	4	2	315	313	-2	10	3	593	595
-9	8	0	144	71	-1	1	1	2641	-2568	3	5	2	1846	1818	5	9	3	241	228
-10	9	0	569	-588	-2	2	1	1505	1535	2	6	2	737	-808	-4	10	3	104	105
-11	10	0	98	-108	-1	3	1	1249	-1210	1	7	2	519	-594	-2	12	3	131	-135
-12	11	0	146	-155	-2	4	1	951	-914	0	8	2	432	-409	-11	13	3	163	189
-13	12	0	102	-102	-3	5	1	401	355	-1	9	2	342	-361	-10	14	3	216	199
-14	13	0	237	232	-4	6	1	123	60	-2	10	2	446	-467	-1	15	3	202	-200
-15	14	0	448	-470	-5	7	1	629	668	1	11	2	730	-761	-10	16	3	474	-462
-16	15	0	190	188	-6	8	1	162	36	5	1	2	286	-303	9	3	3	120	138
-17	16	0	418	-428	-7	9	1	186	194	4	2	2	890	925	8	4	3	306	-317
-18	17	0	243	-235	-8	10	1	337	305	3	3	2	3203	-3415	6	6	3	419	-420
-19	18	0	242	-256	-9	1	1	1967	-2029	2	4	2	173	198	5	7	3	306	303
-20	19	0	135	163	-2	2	1	649	-770	0	6	2	226	239	4	8	3	573	-587
-21	20	0	896	797	-3	3	1	443	423	-1	7	2	109	-174	-3	9	3	130	151
-22	21	0	312	-348	-4	4	1	724	-695	-2	8	2	315	313	-2	10	3	593	595
-23	22	0	245	-250	-5	6	1	470	498	-3	9	2	135	124	-11	13	3	197	149
-24	23	0	224	251	-6	7	1	266	241	-4	10	2	408	-400	-2	14	3	158	162
-25	24	0	145	-145	-7	8	1	580	580	-5	11	2	315	320	-3	15	3	156	-190
-26	25	0	209	-178	-8	9	1	165	139	4	9	1	544	-452	-9</				

Table 4. Continued.

3	7	6	171	-145	14	2	7	196	-180	-11	4	7	301	351	-5	1	8	213	238	6	3	10	264	277	-2	8	1	137	-146	3	1	13	174	-174	
2	7	6	182	-138	12	4	7	400	466	-11	1	7	174	212	-6	2	8	361	-310	5	9	10	221	-225	-4	10	11	166	107	0	4	13	246	-221	
1	9	6	279	-239	11	5	7	232	-203	-12	2	7	162	-182	-7	3	8	130	125	5	11	10	140	-49	3	1	1	269	246	-2	5	13	250	203	
-1	1	6	239	206	7	9	7	163	-147	-13	1	7	159	207	-3	4	8	255	-242	9	3	10	245	-242	1	3	11	106	-17	1	1	13	178	135	
-2	12	6	248	-253	13	1	7	177	132	-11	1	0	108	-111	-9	5	8	162	-146	7	5	10	310	490	0	4	11	134	171	1	1	13	343	-250	
-13	6	182	156	10	4	6	7	346	372	15	3	8	177	26	-10	6	8	262	225	7	5	10	310	490	0	4	11	134	171	1	1	13	343	-250	
6	0	6	908	-955	8	6	7	230	-178	11	5	8	196	-181	-4	0	8	395	390	6	6	10	152	-172	-2	6	11	163	-178	-2	4	13	533	-1076	
7	1	6	204	-259	7	7	7	313	308	6	2	8	157	110	-7	1	8	297	306	4	4	8	132	142	4	6	10	111	131	131	-4	6	13	173	107
6	2	6	121	90	5	9	7	475	-500	4	12	8	204	151	-1	2	8	159	-159	5	7	10	163	-120	-4	4	8	1	197	-173	-3	5	13	193	107
5	3	6	447	457	4	10	7	204	216	2	14	8	163	-109	-9	3	0	503	594	3	9	10	143	-74	-2	4	11	941	959	-2	1	13	255	-256	
4	4	6	263	-239	2	12	7	137	-241	-4	0	8	193	197	-11	5	8	376	-342	2	10	10	179	-134	-3	5	11	149	-138	-4	5	13	265	-261	
3	5	6	452	409	1	13	7	152	162	11	3	8	247	269	-1	0	8	722	720	8	2	10	136	103	-5	7	11	145	-139	-6	5	13	193	107	
2	6	6	223	224	9	3	7	129	-224	-9	5	8	139	-192	-1	3	8	360	330	3	9	10	143	-74	-2	4	11	941	959	-2	1	13	255	-256	
1	7	6	292	200	8	4	7	136	88	8	6	8	151	142	-10	0	8	305	306	6	4	10	202	200	-2	2	11	107	-159	-6	5	13	199	133	
-3	6	6	336	-339	5	7	7	158	190	6	8	8	231	-294	-1	1	8	133	-32	1	9	10	119	41	-4	7	11	571	570	-4	2	13	133	129	
-12	6	177	-164	2	9	7	7	204	-263	5	9	8	272	275	-22	0	8	243	203	1	9	10	132	53	-7	7	11	146	166	-5	3	13	21	-74	
-16	4	6	306	215	0	12	7	312	-292	3	11	8	143	-149	15	1	9	161	-116	0	10	10	177	-191	-7	7	11	146	166	-5	3	13	21	-74	
6	5	6	794	-714	-3	15	7	171	-69	0	11	8	167	-163	-12	1	9	162	-119	-1	11	10	159	177	-4	2	11	237	-249	-8	2	13	197	-189	
4	2	0	122	-156	8	2	7	214	-107	7	5	8	102	-168	10	4	9	220	-170	-2	12	10	142	-42	-5	3	11	175	177	-9	3	13	187	80	
3	3	6	294	-279	6	4	7	222	-237	6	6	8	265	169	5	10	9	240	-172	6	2	10	146	175	-5	1	11	479	479	-7	1	13	194	-120	
2	4	6	395	366	4	6	7	271	-286	5	7	8	140	190	9	3	9	135	81	4	4	10	241	262	-6	2	11	184	-209	7	1	14	111	-75	
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-1	7	6	357	375	-1	11	7	140	-160	1	11	8	186	-244	4	7	9	350	-406	6	0	10	631	-656	-8	2	11	129	98	2	4	14	229	170	
-2	8	6	819	814	7	1	7	356	341	9	1	8	134	171	4	8	9	293	232	5	1	10	381	352	-9	1	11	223	-374	-2	4	14	133	-93	
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-5	1	6	150	-100	5	3	7	166	163	7	5	8	127	-139	0	12	9	193	240	-1	1	10	165	-148	1	5	10	378	376	9	5	12	223	-200	
-7	13	6	177	-167	4	4	7	103	-105	4	3	8	130	135	-1	13	9	155	160	0	6	10	171	175	7	7	12	134	-22	4	2	14	169	35	
-8	4	6	361	196	2	6	7	349	-335	1	9	8	151	-211	5	5	6	9	253	221	-1	7	10	218	-172	9	3	12	226	-222	2	4	14	163	104
1	3	6	1573	-1545	0	8	7	975	-612	-2	12	8	159	97	7	6	9	377	-396	4	4	10	961	-905	2	4	10	122	135	-174	-4	1	14	137	104
-4	0	6	415	414	-1	9	7	231	-229	7	1	8	253	243	2	8	9	266	310	3	1	10	125	105	5	1	10	162	142	-11	0	14	177	62	
-4	5	6	1106	1107	-2	10	7	486	-478	4	8	8	140	-118	7	7	8	249	-222	6	2	10	142	-114	-8	2	12	142	-114	2	2	14	100	84	
-3	6	6	247	-217	1	7	7	239	-169	4	4	8	306	-353	3	5	9	151	-211	5	2	10	146	-93	-3	3	11	147	146	-2	2	14	155	-56	
-4	8	6	762	733	5	1	7	651	634	3	5	8	223	-256	7	6	2	9	133	143	1	3	10	100	-53	7	8	12	112	-59	0	4	14	173	80
-6	10	6	353	-366	4	2	7	240	-210	2	6	8	323	367	5	3	9	234	-230	-1	5	10	279	300	5	5	12	112	-59	0	4	14	173	80	
-12	6	159	-197	2	4	7	148	-146	-2	14	8	143	157	-1	7	9	124	-146	-2	6	10	124	247	-2	6	12	142	-178	-1	1	14	164	140		
4	0	6	120	129	2	4	7	1074	1068	0	8	8	24	253	3	3	9	175	141	-1	1	10	289	-263	-3	1	12	131	-127	-2	6	14	164	-128	
1	1	6	279	-278	1	5	7	333	333	-1	9	8	173	161	-1	9	9	175	171	2	2	10	07	-109	-1	2	12	131	-127	-2	6	14	164	-128	
0	2	6	571	559	0	6	7	143	-131	-5	13	8	182	111	5	1	9	179	-133	1	1	10	289	-263	-3	1	12	131	-127	-2	6	14	164	-128	
-1	4	6	702	-697	3	1	7	436	400	3	3	8	143	157	-1	7	9	124	-146	-2	6	10	124	247	-2	6	12	142	-178	-1	1	14	164	-128	
-2	4	6	948	914	-3	9	7	511	-511	6	0	8	730	725	3	2	9	132	-179	-2	1	10	239	-263	-3	1	12	131	-127	-2	6	14	164	-128	
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-7	9	6	135	125	5	1	7	436	400	3	3	8	143	157	-1	7	9	124	-146	-2	6	10	124	247	-2	6	12	142	-178	-1	1	14	164	-128	
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-2	2	6	156	155	-2	6	7	307	-316	-3	9	8	215	240	-0	4	9	506	506	-5	5	10	241	236	-6	6	10	142	-142	-2	2	14	166	115	
-4	4	6	469	490	-3	7	7	304	314	-4	10	8	170	174	-5	9	9	163	167	-6	6	10	172	-264	0	2	12	152	-119	-3	3	14	207	-150	
-10	6	6	184	187	-1	4	7	201	205	-2	6	8	172	103	-4	0	9	725	-753	-3	3	10	169	-168	-5	3	12	152	-119	-3	3	14	207	-150	
-6	6	6	512	491	-5	9	7	299	-316	4	0	8	1372	1349	-2	1																			

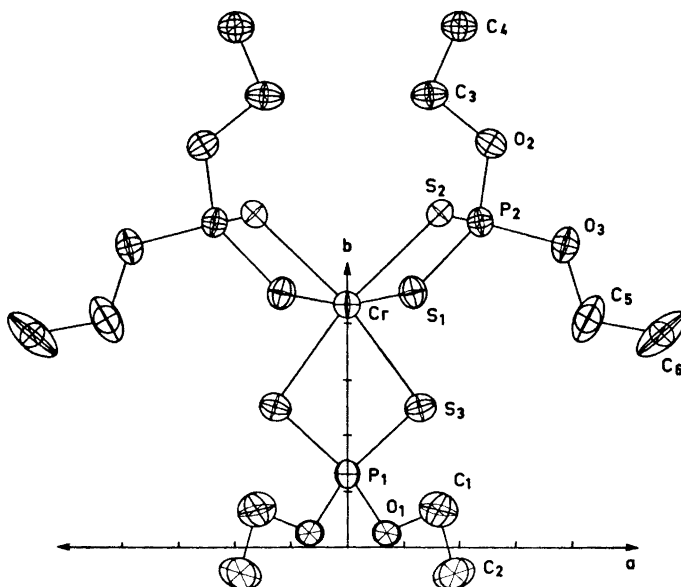


Fig. 1. A molecule of Crdtp_3 projected along the c -axis showing the ellipsoids of vibration.

suggested for Indtp_3 ⁴ some temperature factor parameters could decrease during the refinement (possibly giving less reasonable directions of largest movement) but correlation between atoms which are in space group $C2/c$ related by the twofold axis would be so large that it would be difficult to test whether the deviation from this symmetry was significant. For this reason, refinement in the acentric space group has not been attempted.

The photographs of the other crystalline form were of such poor quality that we could not hope to get reasonably accurate dimensions of the complex in this modification. The dimensions and geometry of the vanadium complex as described in Ref. 3 are so similar to the results found for the chromium complex in this investigation that the only difference has to be found in the packing of the molecules.

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Received July 21, 1971.